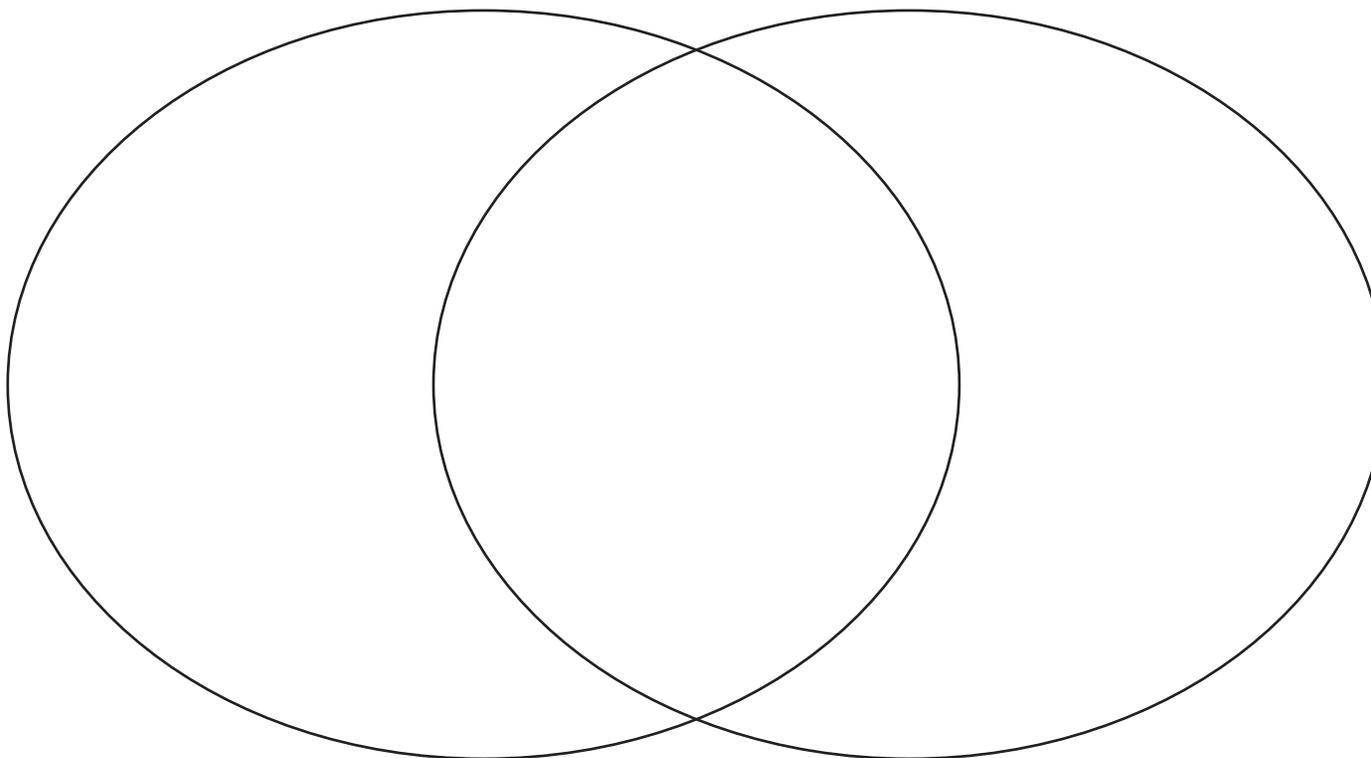
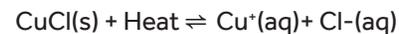
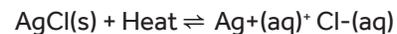


# EQUILIBRIUM

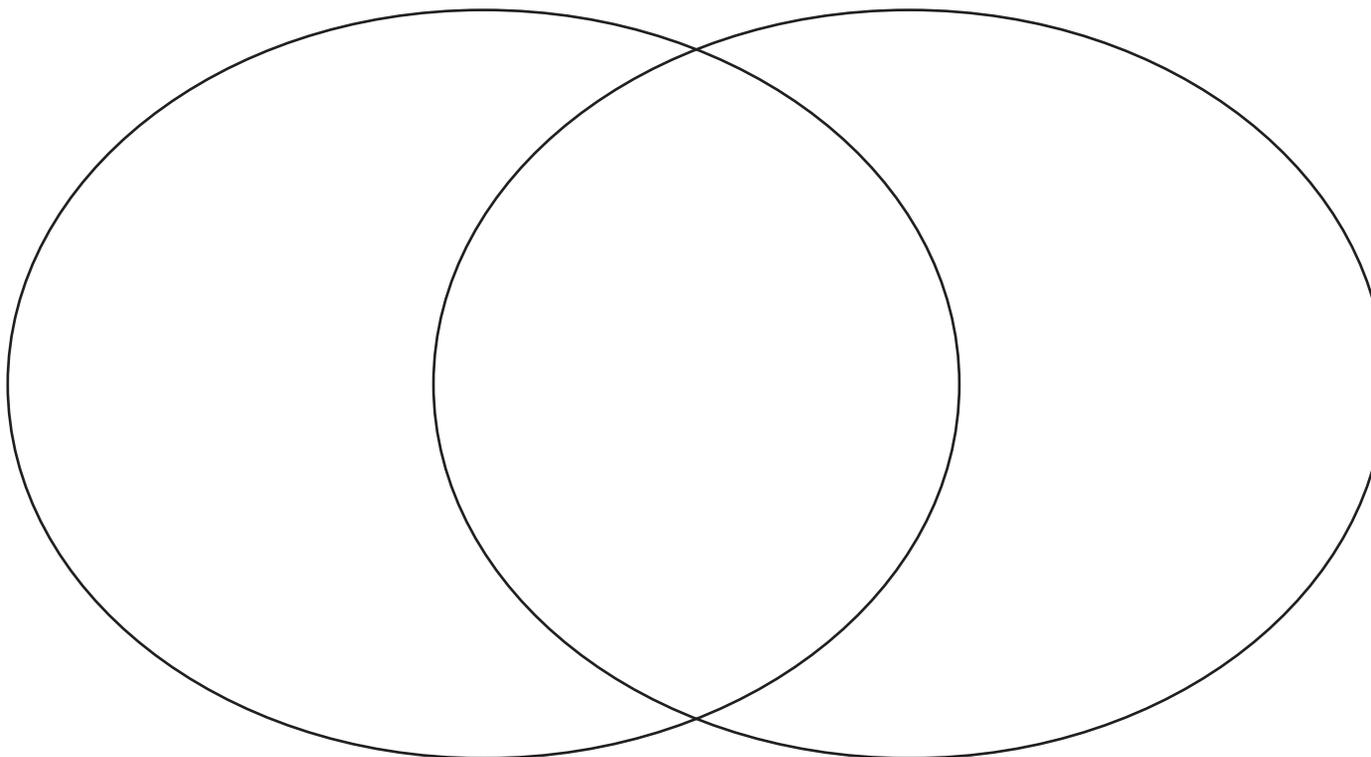
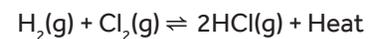
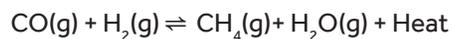
## Directions:

1. Log into Collisions and navigate to the Equilibrium Game.
2. Play the Tutorial levels, if you haven't done so already.
3. Exit the levels and enter the Equilibrium Sandbox.
4. Use the Sandbox to fill in the Venn diagrams using the statements in the bank beneath it. **NOT ALL STATEMENTS SHOULD BE USED!**



Removing $\text{Ag}^+$ ions would shift the reaction left.	Heat can be thought of as a reactant.	Addition of $\text{AgCl(aq)}$ solution would shift the reaction to the left.	Decreasing the temperature would increase the presence of the solid.	Endothermic
Addition of $\text{Cu}^+$ would shift the reaction to the left.	Increasing the temperature produces more aqueous ions.	Changing the pressure has no effect on the equilibrium.	Adding $\text{Ag}^+$ ions would shift the reaction left.	Exothermic

# EQUILIBRIUM



Decrease in pressure shifts the reaction to the left.	Respond(s) to an increase in pressure by shifting towards the products.	Increase in CO concentration shifts the reaction to the right.	Decrease in Cl <sub>2</sub> concentration shifts reaction to the left.	Endothermic
Increase in HCl concentration shifts the reaction to the right.	Respond(s) to an increase in temperature by shifting to the left.	Decrease in CH <sub>4</sub> concentration shifts the reaction to the right.	Decreasing the pressure of this reaction would have no effect on the equilibrium.	Exothermic
Respond(s) to a decrease in temperature by shifting to the left.	Addition of H <sub>2</sub> would shift the reaction to the right.	Increase in pressure shifts the reaction to the right.	If the product were removed as it were created, the reaction would shift to the right.	